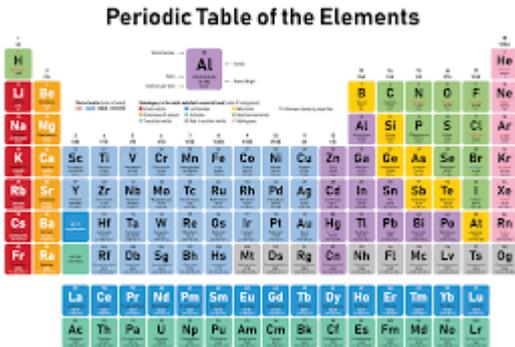
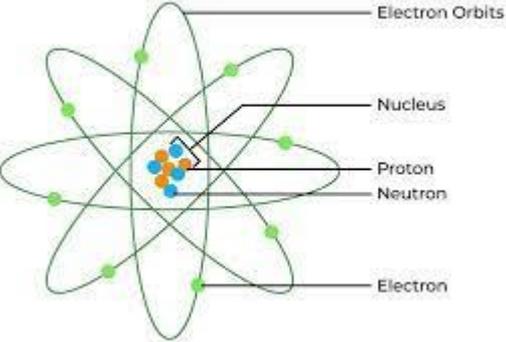
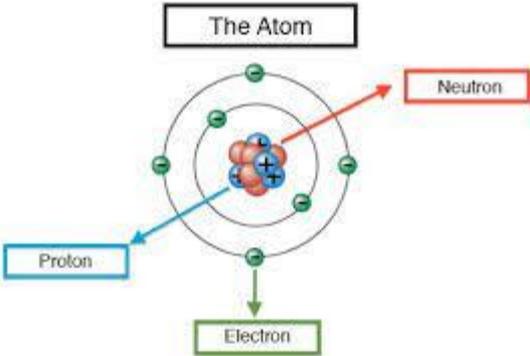


# BASIC 8

## WEEKLY LESSON PLAN – WEEK 5

<b>Strand:</b>	Diversity of Matter	<b>Sub-Strand:</b>	Materials
<b>Content Standard:</b>	B8.1.2.2 Demonstrate understanding of atoms and the atomic structure of elements in the periodic table		
<b>Indicator (s)</b>	B8.1.2.2.1 Describe atoms as composed of sub-atomic particles	<b>Performance Indicator:</b> Learners can identify the location of sub-atomic particles in an atom	
<b>Week Ending</b>	10-02-2023		
<b>Class</b>	B.S.8	<b>Class Size:</b>	<b>Duration:</b>
<b>Subject</b>	Science		
<b>Reference</b>	Science Curriculum, Teachers Resource Pack, Learners Resource Pack, Textbook.		
<b>Teaching / Learning Resources</b>	Bottle tops, salt, sugar, sand, gari, gravel, oil, water, Poster, Pictures.	<b>Core Competencies:</b>	
<b>DAY/DATE</b>	<b>PHASE 1 : STARTER</b>	<b>PHASE 2: MAIN</b>	<b>PHASE 3: REFLECTION</b>
<b>MONDAY</b>	<p>Write keywords and terminologies on the chalkboard.</p> <p>Assist Learners to explain the meanings of the terminologies.</p>	<ol style="list-style-type: none"> <li>1. Show Learners picture displaying the periodic table.</li> <li>2. Learners to observe the periodic table.</li> <li>3. Assist Learners to explain Atom and structure of an element using the periodic table.</li> </ol> <div style="text-align: center;">  <p><b>Periodic Table of the Elements</b></p> </div>	<p>Through questions and answers, conclude the lesson.</p> <p><b>Exercise;</b></p> <p>Draw the periodic table.</p>

		<p style="text-align: center;"><b>Lewis Structures of All Elements</b></p> <div style="text-align: center;"> <table border="1" style="margin: auto;"> <tr> <td style="padding: 5px;">·H</td> <td style="padding: 5px;">:He</td> <td style="padding: 5px;">:Ḃ</td> <td style="padding: 5px;">·Ċ·</td> </tr> <tr> <td style="padding: 5px;">·Ṅ·</td> <td style="padding: 5px;">:Ȯ·</td> <td style="padding: 5px;">:Ḟ·</td> <td style="padding: 5px;">:Nė:</td> </tr> </table> </div> <div style="text-align: center; margin-top: 20px;">  </div>	·H	:He	:Ḃ	·Ċ·	·Ṅ·	:Ȯ·	:Ḟ·	:Nė:	
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<p><b>THURSDAY</b></p>	<p>Discuss the meanings of proton, electron and neutron</p>	<ol style="list-style-type: none"> <li>1. Assist Learners to identify the sub-atomic particles found in an atom.</li> <li>2. Discuss the locations of Sub-atomic particles in an atom.</li> <li>3. Learners brainstorm to draw the structure of an atom indicating the locations of the sub-atomic particles.</li> </ol> <p><b>Sub-Atomic Particles in an atom;</b></p> <ul style="list-style-type: none"> <li>● Protons</li> <li>● neutrons</li> <li>● Electrons</li> </ul> <div style="text-align: center; margin-top: 20px;">  </div>	<p>Individual Learners to drawing the structure atom and show the location of the sub-atomic particles</p>								
<p><b>FRIDAY</b></p>	<p>Review Learners knowledge on the previous lesson.</p>	<ol style="list-style-type: none"> <li>1. Assist Learners to explain the meaning of electrical charges.</li> <li>2. Discuss with Learners the electrical charges of the sub-atomic particles</li> <li>3. Demonstrate calculating for the number of sub-atomic particles using formula</li> <li>4. Learners in small groups to calculate for the number of electrical charges</li> </ol> <p>Atomic number = the number of protons</p>	<p>Through questions and answers, conclude lesson.</p>								

		<p>Mass number = the number of protons + the number of neutrons From these, we can deduct:</p> <p>Number of neutrons = Mass number – Atomic number</p> <p>Atom has no overall charge, which means there are equal number of negatively charged electrons and positively charged protons. If we know the number of protons (or atomic number) of an atom, this will be equal to the number of electrons of that atom. Number of electrons = Number of protons</p>	
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**Name of Teacher:**

**School:**

**District:**